## Chemistry 202L
## Practicum

**Name:_____Bromination_____**



$$\text{C}_{13}\text{H}_{11}\text{NO} \quad + \quad \text{Br}_2 \quad \xrightarrow[\text{conc HBr}]{\text{CH}_3\text{CO}_2\text{H}} \quad \text{Product ?}$$

Mol. Wt.
Volume    solid
Density    -----
%          100%
Grams      .098g
Moles
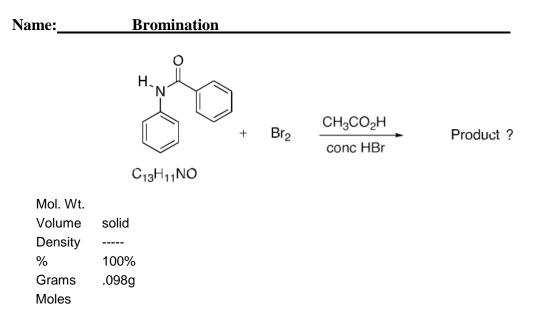
1. Using your laboratory textbook and your write-up of a similar experiment, **outline a pre-lab**: grams of reactants and solvents, reaction conditions, quenching, work- up, purification, analysis.
2. Bromine comes as a 24% by weight bromine solution in conc HBr (density=1.70g/mL).
3. You will need to use 1.5 equivalents of Br₂.
4. You will need to heat the benzanilide/acetic acid mixture in a 55-60° C water bath until the benzanilide dissolves.
5. You will then turn off the heat and remove the water bath. The Br₂ solution will then be added.